

## Effects of ion-exchange for NOM removal in water treatment with ceramic membranes ultrafiltration

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**Abstract.** To enhance the efficiency of water treatment and reduce the extent of membrane fouling, the membrane separation process is frequently preceded by other physico-chemical processes. One of them might be ion exchange. The aim of this work was to compare the efficiency of natural organic matter removal achieved with various anion-exchange resins, and to verify their potential use in water treatment prior to the ultrafiltration process involving a ceramic membrane. The use of ion exchange prior to ceramic membrane ultrafiltration enhanced final water quality. The most effective was MIEX, which removed significant amounts of the VHA, SHA and CHA fractions. Separation of uncharged fractions was poor with all the resins examined. Water pretreatment involving an ion-exchange resin failed to reduce membrane fouling, which was higher than that observed in unpretreated water. This finding is to be attributed to the uncharged NOM fractions and small resin particles that persisted in the water.

**Keywords:** natural organic matter; ion-exchange resin; pretreatment; ceramic membrane

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### 1. Introduction

Ultrafiltration (UF) is an advanced treatment technology providing efficient removal of a broad spectrum of pollutants from the water being treated. That is why the application of this membrane process in water treatment plants has become increasingly frequent. Among the diversity of pollutants that ought to be removed, natural organic matter (NOM) deserves particular attention. Natural dissolved matter removal is an essential part of the treatment process, principally because this offers the possibility to reduce the risk that disinfection by-products may form during water chlorination (Halim *et al.* 2010). Another reason behind the necessity of NOM removal from the water being treated is the contribution of these organic species to the increase in the colour intensity of the water and the associated need to use much higher coagulant or disinfectant doses than those required for unpolluted water treatment.

Research reported in the literature has revealed that owing to their relatively large pores UF membranes are able to remove only large NOM fractions (Schäfer 2001). It has furthermore been demonstrated that natural organic matter is responsible for membrane fouling (Lee *et al.* 2004, Raspati *et al.* 2011, Yeh *et al.* 2012). Thus, in order to enhance the efficiency of NOM separation and reduce the extent of membrane fouling, ultrafiltration is frequently combined with other physico-chemical processes. Among them preference is given to coagulation and active carbon

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adsorption, which are used for NOM removal prior to membrane filtration. However, in some instances such pretreatment fails to provide sufficient removal of organic matter fractions responsible for membrane blocking, and thus raises the treatment cost. These findings have spurred the search for more efficient methods. One of them might be the ion-exchange process.

Many NOM molecules contain carboxylic acid structures, which make them anionic in nature and removable by anion-exchange resins (AER). The removal of NOM *via* anion exchange is greatly influenced by the properties of the resin applied. Bolto *et al.* (2002) reported that weak-alkaline resins remove less NOM as compared with strongly alkaline ones, and that resins with a macroporous structure remove larger quantities of NOM than do gel-type resins. The latter result, however, contradicts the findings reported by Tan *et al.* (2005), who obtained higher NOM removal with gel-type resins. It is essential to note that the higher NOM removal observed in the case of gel-type resins should be attributed to their higher swelling capacity in water. Another contributory factor in NOM removal is the material from which the resins have been made. Gottlieb (1996) demonstrated that resins with a styrene structure display a greater affinity with aromatic components than do resins based on an acrylic skeleton.

New opportunities for the application of ion exchange to NOM removal have been created since Orica Watercare developed the MIEX<sup>®</sup>DOC process, which makes use of a magnetic ion-exchange resin (MIEX). Its grains, made of strong-base anion-exchange resin, are 2–5 times smaller than conventional AER beads, which provides a greater external surface area and thus allows rapid sorption kinetics. Moreover, the magnetic iron oxide incorporated in the resin structure causes the beads to agglomerate together over time allowing for easy removal once the treatment is complete.

The use of the combined process ion exchange/low pressure membrane filtration has been the subject of several laboratory and pilot-scale studies. Cornelissen *et al.* (2009, 2010), who analyzed the influence of fluidized ion-exchange pretreatment on the performance of ultra- and nanofiltration membranes, noticed that the decrease in NOM concentration, even though significant, was not paralleled by a decline in membrane fouling. Li *et al.* (2009) observed a decrease in the extent of UF membrane fouling, when the integrated process was carried out with a weak cation-exchange resin. This improvement in the transport properties of the membrane was attributable to the removal of calcium ions.

In recent years focus has also been placed on the MIEX<sup>®</sup>DOC process when used as a pretreatment step to MF or UF (Fabris *et al.* 2007, Zhang *et al.* 2007, Kabsch-Korbutowicz *et al.* 2006, Humbert *et al.* 2007, Dixon *et al.* 2010, Drikas *et al.* 2011). All of the investigators arrived at the conclusion that although those integrated processes produced high quality water, their effect on the extent of membrane fouling was discouraging, as it largely depended on the properties of the membrane used (e.g., pore size) and the composition of the water being treated (e.g., NOM properties or Ca<sup>2+</sup> concentration).

Despite the importance attached to the integration of different treatment technologies in order to improve the quality of finished water, no references have been found in the literature to the application of the anion-exchange process as a prior step to ultrafiltration carried out with ceramic membranes. At present, in the sector of water treatment preference is given to the use of polymeric membranes. Nevertheless, owing to the excellent separation properties of inorganic membranes, as well as to the sharp decrease in the cost of their manufacture, the use of ceramic membranes in water treatment plants shows a tendency to increase.

The aim of this work was to compare the efficiency of natural organic matter removal achieved with various anion-exchange resins, and to verify their potential use in water treatment prior to the

ultrafiltration process involving a ceramic membrane.

## 2. Materials and methods

### 2.1 Feed solution

The experiments were carried out using a model solution prepared from dechlorinated tap water and humic-rich water flowing out from a peat-bog in the Table Mountains (Poland) (sampling point 50°27'29.97"N; 16°23'16.87"E). NOM concentration was monitored by measuring DOC, UV absorbance at 254 nm, and colour intensity (Shimadzu UV1240 spectrophotometer). The properties of the feed solutions are shown in Table 1.

The efficiency of NOM separation was analyzed in terms of change in retention coefficient (R).

### 2.2 Ion exchange

For the purpose of this study use was made of five anion-exchange resins. Their characteristics are specified in Table 2. The resins differed in skeleton matrix, ion-exchange capacity and particle diameter. The experimental resin doses amounted to 2.5 cm<sup>3</sup>/dm<sup>3</sup>. An appropriate amount of the resin was dosed to 1 dm<sup>3</sup> of treated water, and then the resin-enriched solution was stirred at 135 rpm for 20 min. After that the samples were left for sedimentation for 30 min.

Table 1 Properties of feed solution

Parameter	Range	Average
Colour, g Pt/m <sup>3</sup>	51.7 – 53.1	52.3
UV absorbance at 254 nm, cm <sup>-1</sup>	0.339 – 0.346	0.339
Dissolved organic carbon, g C/m <sup>3</sup>	7.9 – 8.3	8.1
pH	7.3 – 7.4	-
Conductivity, μS/cm	720 – 770	735

Table 2 Anion-exchange resin characteristics

Resin name	Resin type	Polymer structure	Ion-exchange capacity, mmol/cm <sup>3</sup>	Water content, %	Particle diameter, mm
Purolite A100	weak base	macroporous polystyrene	0.885	53 – 60	0.60 – 0.85
Purolite A200	strong base type II	gel polystyrene crosslinked with divinylbenzene	1.040	45 – 54	0.60 – 0.85
Purolite A400	strong base type I	gel polystyrene crosslinked with divinylbenzene	1.040	48 – 54	0.60 – 0.85
Wofatit SBW	strong base type I	sulphonated gel polystyrene crosslinked with divinylbenzene	0.731	45 – 48	0.30 – 1.20
MIEX <sup>®</sup>	strong base type I	macroporous polyacrylate	0.400	65	0.15 – 0.18



Fig. 1 Ultrafiltration cross-flow system

### 2.3 Ultrafiltration

Ultrafiltration experiments were performed with a 1-channel 15 kDa ceramic  $ZrO_2/TiO_2$  membrane made by Tami Industries. Channel diameter and effective filtration surface amounted to 6 mm and 40  $cm^2$ , respectively. The cross-flow pilot set-up used in the experiments was delivered by J.A.M. INOX PRODUKT (Fig.1) and consisted of a reservoir tank (10  $dm^3$ ), a pump, pressure gauges, a membrane module and a flowmeter for retentate. Both retentate and permeate were recirculated to the stirred feed tank in order to enable steady-state operation. The system was thermostated and water temperature was kept at 20°C. The process was run at a pressure of 0.3 MPa.

### 2.4 Integrated process

The integrated ion exchange/ultrafiltration process was carried out using a decanted water sample, which after treatment in the ion-exchange process, followed by 30-min sedimentation, was ultrafiltered at a pressure of 0.3 MPa.

### 2.5 NOM fractionation

The NOM fractionation procedure was performed according to the method described by Chow *et al.* (2004). Prior to fractionation, the water samples were pre-filtered using a 0.45  $\mu m$  membrane. A 500  $cm^3$  volume of water adjusted to pH 2 was passed through a column packed with 15  $cm^3$  DAX-8 resin to adsorb very hydrophobic acids (VHA). The effluent from the first column was passed

through the column packed with 15 cm<sup>3</sup> XED-4 resin to adsorb slightly hydrophobic acids (SHA). The effluent from the second column was alkalized to pH 8 and passed through the column containing 15 cm<sup>3</sup> IRA-958 resin to adsorb charged hydrophilic substances (CHA). The effluent from the third column contained neutral hydrophilic substances (NEU).

The concentrations of particular NOM fractions were calculated according to the following equations

$$\begin{aligned} \text{VHA} &= \text{DOC}_{\text{raw}} - \text{DOC}_{\text{DAX-8 effluent}} \\ \text{SHA} &= \text{DOC}_{\text{DAX-8 effluent}} - \text{DOC}_{\text{XAD-4 effluent}} \\ \text{CHA} &= \text{DOC}_{\text{XAD-4 effluent}} - \text{DOC}_{\text{IRA-985 effluent}} \\ \text{NEU} &= \text{DOC}_{\text{IRA-985 effluent}} \end{aligned}$$

### 3. Results and discussion

The efficiency of NOM removal obtained with the resins tested is depicted in Fig. 2. The best water purification effect, i.e. the highest NOM removal, was achieved with the MIEX resin. Colour intensity was reduced by 54.2%, and DOC retention reached 36.7%, although for the other resins examined, regardless of their properties, it did not exceed 10%. The positive treatment effects achieved with MIEX can be attributed to the macroporous structure and the high water content of this resin. Moreover, smaller MIEX particles provide greater external surface areas. Under such conditions, organic anions can easily diffuse into the pores of the resin.

Analysis of the amount of the residual resin that persisted in the water after 30-minute sedimentation (Fig. 3) suggests that some fine particles remaining in the water may affect the quality of the stream that feeds the UF membrane. Surprisingly, the highest residual resin concentration was measured when the process was carried out with MIEX. This resin contains magnetic elements which should accelerate the formation of easily settleable aggregates. Probably, due to multiple use and regeneration of the resin, some of its beads pulverized, and thus deteriorated the efficiency of their removal during sedimentation.

Our previous research (Kabsch-Korbutowicz *et al.* 2009) has established the following: ultrafiltration

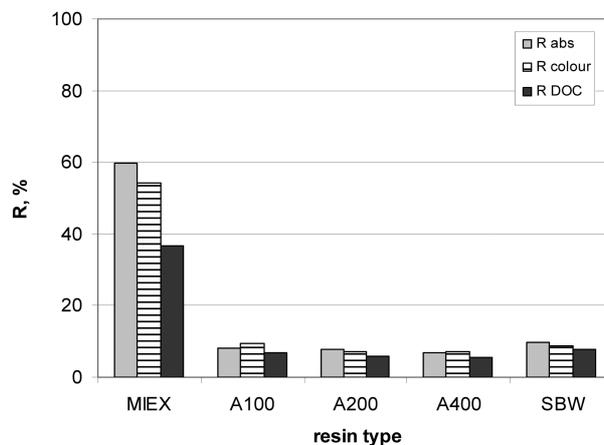


Fig. 2 Influence of resin type on NOM removal

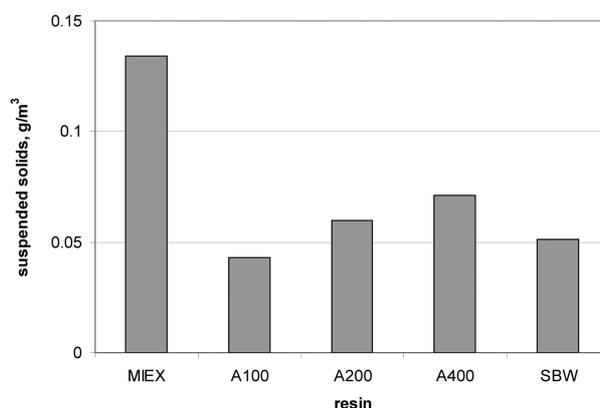


Fig. 3 Concentration of residual resin particles after 30-minute sedimentation

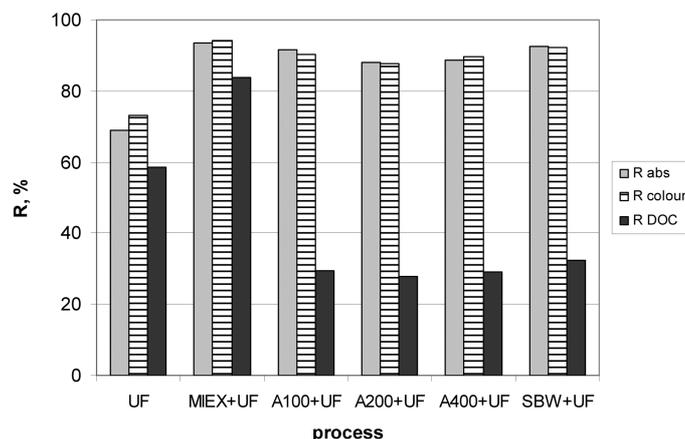


Fig. 4 Efficiency of NOM removal in the integrated ion exchange/UF process

with ceramic membranes as a sole process fails to provide sufficiently high NOM removal, and the extent of membrane fouling still remains high. In the study reported on in this paper, integration of the ultrafiltration process with ion exchange as a prior step enhanced NOM separation for all the membranes tested, the best treatment effects being achieved with the MIEX resin (Fig. 4). Combination of the MIEX<sup>®</sup>DOC process with UF accounted for an approximately 84% reduction in DOC and a 94% reduction in colour intensity. The use of the other resins also improved NOM separation, but the improvement was not as spectacular as in the case of MIEX. With all the resins tested, the retention coefficient values achieved for UV absorbance and colour intensity were noticeably higher than those obtained for DOC. Water colour and UV absorbance indicate the quantity of large NOM fractions containing, among others, aromatic structures, whereas DOC concentration describes the amount of all organic substances in the water being treated.

Apart from enhancing the efficiency of NOM separation, the integration of ion exchange with the UF process is intended for reducing the extent of membrane fouling. In our study, however, such integration failed to sufficiently reduce the fouling effect, when the solution was pretreated by ion

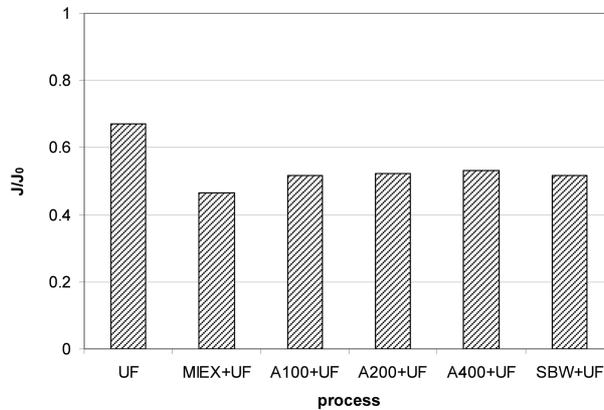


Fig. 5 Influence of water pretreatment on relative permeate flux

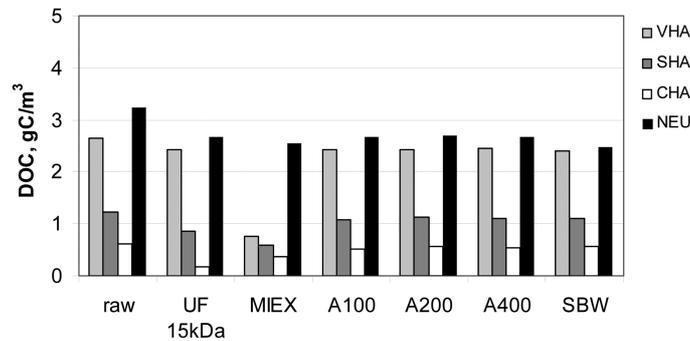


Fig. 6 Proportion of NOM fraction in raw solution, permeate of 15 kDa membrane and samples after ion exchange

exchange before being passed to the process of ultrafiltration on the ceramic 15 kDa membrane. Fig. 5 shows the contribution of ion-exchange pretreatment to the decline in permeate flux. The normalized permeate flux ( $J/J_0$ ) represents the flux of the model solution ( $J$ ), with or without pretreatment, compared to the clean water flux ( $J_0$ ) of the membrane using redistilled water. Without pretreatment,  $J/J_0$  declined to approximately 0.67. When the solution was subject to anion-exchange pretreatment before ultrafiltration, the normalized permeate flux ranged between 0.46 and 0.53. The most significant decline in permeate flux was observed when the MIEX resin was used.

The explanation lies in the composition of the fractions removed via ion exchange. As shown in Fig. 6, the raw solution was composed of very hydrophobic acids (34.2% of total DOC), slightly hydrophobic acids (15.9%), charged hydrophilic substances (7.9%), and hydrophilic neutrals (42%). We observed that all of the charged compounds (very hydrophobic, slightly hydrophobic, and charged hydrophilics) were removed with high efficiency in the MIEX process, whereas the neutral fractions remained unaffected. With the other resins examined, the efficiency of charged species separation was very low. We may assume that it is the uncharged hydrophilic fraction of NOM that should be blamed for membrane fouling. Such assumption is consistent with the findings reported by other researchers (Cho *et al.* 2000, Kimura *et al.* 2004) who suggested that larger-sized hydrophilic

neutrals (e.g. polysaccharide-like organics) are the strongest foulants. Cornelissen *et al.* (2010) and Hofs *et al.* (2012) observed that although pretreatment with AER provided removal of humic substances and hydrophobic organic carbon, it failed to produce the removal of biopolymers, building blocs or neutrals. Fouling of membranes may also result from the presence of colloids that may still be present in water after ion-exchange. Seemingly, it is also the presence of residual resin in the water after pretreatment that may intensify membrane blocking. Fine resin particles that occur in the water may plug the pores of, or form a compact cake on, the membrane and thus increase the fouling effect. We are not ruling out the possibility that the interactions between fine resin particles and NOM may also be regarded as contributory factors in the intensification of membrane blocking.

### 3. Conclusions

The use of ion exchange prior to ceramic membrane filtration of water containing natural organic matter enhances the efficiency of the treatment process for NOM removal. Among the resins tested, definitely best results of NOM removal were achieved with the MIEX resin. Despite the removal of natural organic matter by the anion-exchange resin, the hydraulic performance of the ceramic 15 kDa ultrafiltration membrane did not improve. Seemingly, the increase in the extent of membrane fouling is to be attributed to the presence of the neutral NOM fraction and fine resin particles, which persisted in the water after pretreatment.

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### References

- Bolto, B., Dixon, D., Eldridge, R., King, S. and Linge, K. (2002), "Removal of natural organic matter by ion exchange", *Water Res.*, **36**(20), 5057-5065.
- Cho, J., Amy, G. and Pellegrino, J. (2000), "Membrane filtration of natural organic matter: factors and mechanisms affecting rejection and flux decline with charged ultrafiltration (UF) membrane", *J. Membrane Sci.*, **164**(1-2), 89-110.
- Chow, Ch.W.K., Fabris, R. and Driks, M. (2004), "A rapid fractionation to characterize natural organic matter for the optimization of water treatment process", *J. Water Supply: Res. Technol. – AQUA*, **53**(2), 85-92.
- Cornelissen, E.R., Beerendonk, E.F., Nederlof, M.N., van der Hoek, J.P. and Wessels, L.P. (2009), "Fluidized Ion Exchange (FIX) to control NOM fouling in ultrafiltration", *Desalination*, **236**(1-3), 334-341.
- Cornelissen, E.R., Chasseriaud, D., Siegers, W.G., Beerendonk, E.F. and van der Kooij, D. (2010), "Effect of anionic fluidized ion exchange (FIX) pre-treatment on nanofiltration (NF) membrane fouling", *Water Res.*, **44**(10), 3283-3293.
- Dixon, M.B., Morran, J.Y. and Driks, M. (2010), "Extending membrane longevity by using MIEX as pre-treatment" *J. Water Supply: Res. Technol. – AQUA*, **59**(2-3), 92-99.
- Driks, M., Dixon, M. and Morran, J. (2011), "Long term case study of MIEX pre-treatment in drinking water; understanding NOM removal", *Water Res.*, **45**(4), 1539-1548.
- Fabris, R., Lee, E.K., Chow, C.W.K., Chen, V. and Driks, M. (2007), "Pre-treatments to reduce fouling of low

- pressure micro-filtration (MF) membranes”, *J. Membrane Sci.*, **289**(1-2), 231-240.
- Gottlieb, M. (1996), “The reversible removal of naturally occurring organics using resins regenerated with sodium chloride”, *Ultrapure Water*, **11**, 53–58.
- Halim R., Utama R., Cox S. and Le-Clech P. (2010), “Performances of submerged membrane photocatalysis reactor during treatment of humic substances”, *Membrane Water Treatment, An Int. J.*, **1**(4).
- Hofs, B., Vries D., Siegers, W.G., Beerendonk, E.F. and Cornelissen, E.R. (2012), “Influence of water type and pretreatment method on fouling and performance of an Al<sub>2</sub>O<sub>3</sub> microfiltration membrane”, *Desalination*, **299**(1), 28-34.
- Humbert, H., Gallard, H., Jacquement, V. and Croué, J-P. (2007), “Combination of coagulation and ion exchange for the reduction of UF fouling properties of high DOC content surface water”, *Water Res.*, **41**(17), 3803-3811.
- Kabsch-Korbutowicz, M. and Urbanowska, A. (2009), “Application of ceramic membranes to the removal of natural organic matter from water”, *Ochrona Srodowiska*, **31**(1), 15-19 (in Polish).
- Kabsch-Korbutowicz, M., Biłyk, A. and Mołczan, M. (2006), “The effect of feed water pretreatment on ultrafiltration membrane performance”, *Polish J. Environ. Studies*, **15**(5), 719-725.
- Kimura, K., Hane, Y., Watanabe, Y., Amy, G. and Ohkuma, N. (2004), “Irreversible membrane fouling during ultrafiltration of surface water”, *Water Res.*, **38**(14-15), 3431-3441.
- Lee, N., Amy, G., Croué, J-P. and Buisson, H. (2004), “Identification and understanding of fouling in low-pressure membrane (MF/UF) filtration by natural organic matter (NOM)”, *Water Res.*, **38**, 4511–4523.
- Li, S., Heijman, S.G.J., Verberk, J.Q.J.C. and van Dijk, J.C. (2009), “An innovative treatment concept for future drinking water production: fluidized ion exchange – ultrafiltration – nanofiltration granular activated carbon filtration”, *Drinking Water Eng. Sci.*, **1**, 41-47.
- Raspati, G.S., Høvik, H.N. and Leiknes, T. (2011), “Preferential fouling of natural organic matter (NOM) fractions in submerged low-pressure membrane filtration”, *Desalination Water Treatment*, **34**(1-3), 416-422.
- Schäfer, A.I. (2001), *Natural Organics Removal Using Membranes: Principles, Performance and Cost*, Technomic Publishing Company, Inc., Lancaster.
- Tan, Y., Kilduff, J.E., Kitis, M. and Karanfil, T. (2005), “Dissolved organic matter removal and disinfection byproduct formation control using ion exchange”, *Desalination*, **176**(1-3), 189-200.
- Yeh, H.-M., Ho, C.-D. and Li, C.-H. (2012), “Effects of ring number and baffled-ring distances on ultrafiltration in the tubular membrane inserted concentrically with a ring rod”, *Membrane Water Treatment, An Int. J.*, **3**(1), 51-62.
- Zhang, R., Vigneswaran, S., Ngo, H. and Nguyen, H. (2007), “A submerged membrane hybrid system coupled with magnetic ion exchange (MIEX®) and flocculation in wastewater treatment”, *Desalination*, **216**(1-3), 325-333.